

Name: SOLUTIONS

Date: JAN 15/15

Chemistry Review Quiz  
SNC 2P1

K	/25	A	/9
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1. Complete the following table:

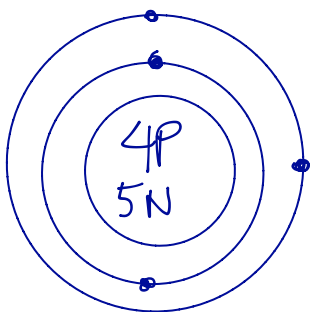
[0.5 K marks each, total of 8]

Symbol	Name	Atomic Number	Atomic Mass	Number of Protons	Number of Neutrons	Number of Electrons
C	Carbon	6	12	6	6	6
Si	silicon	14	28.09 OR 28	14	14	14
Al	Aluminium	13	26.98 OR 27	13	27 - 13 = 14	13

2. Draw the Bohr-Rutherford diagram for the following atoms:

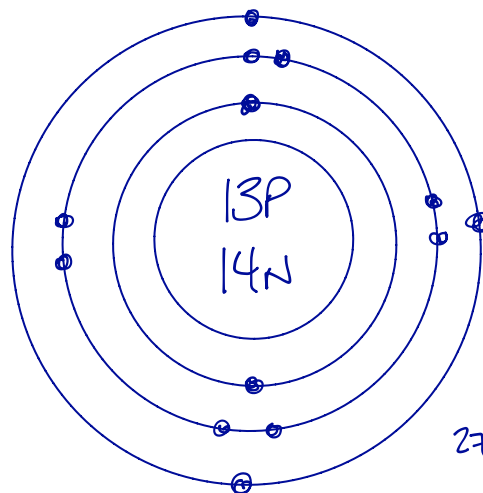
[2 K marks each]

a) Beryllium



$9 - 4 = 5$  NEUTRONS

b) Aluminum



$27 - 13 = 14$  NEUTRONS

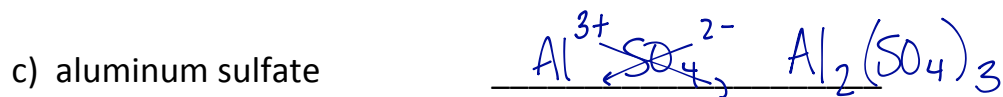
3. Define "ion": CHARGED ATOM/PARTICLE

[1 K mark]

4. Identify if each of the following compounds is ionic or molecular and name it in the table below. [1 K marks each]

	Ionic (I) or Molecular (M)	Compound Name
a) $\text{CaF}_2$	I	calcium fluoride
b) $\text{CO}$	M	carbon monoxide
c) $\text{N}_2\text{O}_5$	M	dinitrogen pentoxide
d) $\text{NaOH}$	I	sodium hydroxide

5. Find the chemical formula of the following compounds: [1 K mark each]



*Hint: polyatomic table can be found on the back of the quiz!*

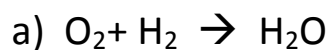
6. Identify each reaction as one of the following types of reaction: [1 A mark each]

Synthesis

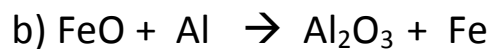
Decomposition

Single displacement

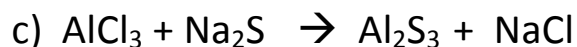
Double Displacement



Reaction type: synthesis

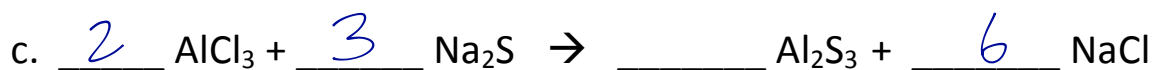
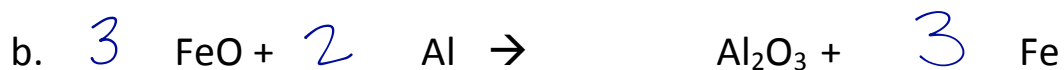


Reaction type: single displacement



Reaction type: double displacement

7. Balance the following chemical equation by filling in the blanks: [2 A marks each]



<b>Name of polyatomic ion</b>	<b>Ion formula</b>	<b>Name of polyatomic ion</b>	<b>Ion formula</b>
Nitrite	$\text{NO}_2^-$	Hydrogen Carbonate (Bicarbonate)	$\text{HCO}_3^-$
Nitrate	$\text{NO}_3^-$		
Sulfite	$\text{SO}_3^{2-}$	Carbonate	$\text{CO}_3^{2-}$
Sulfate	$\text{SO}_4^{2-}$	Phosphate	$\text{PO}_4^{3-}$
Chlorite	$\text{ClO}_2^-$	Acetate	$\text{C}_2\text{H}_3\text{O}_2^-$
Chlorate	$\text{ClO}_3^-$	Cyanide	$\text{CN}^-$
Arsenite	$\text{AsO}_3^{3-}$	Hydroxide	$\text{OH}^-$
Arsenate	$\text{AsO}_4^{3-}$	Ammonium	$\text{NH}_4^+$